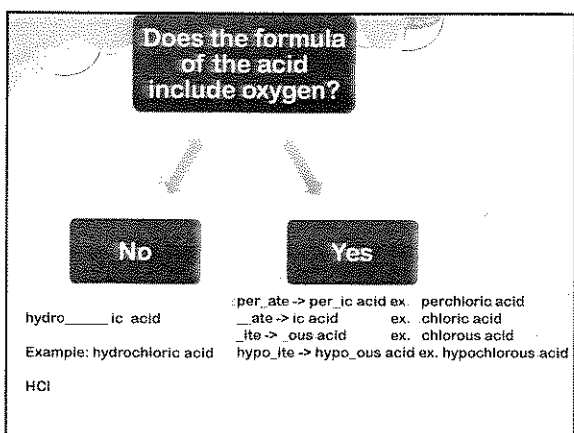


Oxyacids (Acids With Oxygen)		
Acid Name	Chemical Formula	Parent Oxyanion
nitric acid	$\text{HNO}_{3(aq)}$	nitrate, $\text{NO}_3^-$
chloric acid	$\text{HClO}_{3(aq)}$	chlorate, $\text{ClO}_3^-$
carbonic acid	$\text{H}_2\text{CO}_{3(aq)}$	carbonate, $\text{CO}_3^{2-}$
sulfuric acid	$\text{H}_2\text{SO}_{4(aq)}$	sulfate, $\text{SO}_4^{2-}$
phosphoric acid	$\text{H}_3\text{PO}_{4(aq)}$	phosphate, $\text{PO}_4^{2-}$

Oxyanions of Chlorine	
ANION	Formula of ANION
perchlorate	$\text{ClO}_4^-$
chlorate	$\text{ClO}_3^-$
chlorite	$\text{ClO}_2^-$
hypo	$\text{ClO}^-$

How to name oxyacids?	Example
per_ate anion → per_ic acid	perchloric acid
ate anion → ic acid	chloric acid
ite anion → ous acid	chlorous acid
hypo_ous anion → hypo_ous acid	hypochlorous acid



What is the name of the acid  $\text{H}_2\text{SO}_{3(aq)}$ ?

Start by looking at the charts. Is oxygen present? Oxygen IS present so we find the oxyanion that it's related to.

The oxyanion is sulfite  $\text{SO}_3^{2-}$ .

Establish the name of the acid from the name of the oxyanion. ite → ous

sulfurous acid