

1	H	1
2	Li	3 Be 4
3	Na	11 Mg 12
4	K	19 Ca 20
5	Rb	37 Sr 38
6	Cs	55 Ba 56
7	Fr	87 Ra 88

Sc	21	Ti	22	V	23	Cr	24	Mn	25	Fe	26	Co	27	Ni	28	Cu	29	Zn	30	Ga	31	Ge	32	As	33	Se	34	Br	35	Kr	36
Y	39	Zr	40	Nb	41	Mo	42	Tc	43	Ru	44	Rh	45	Pd	46	Ag	47	Cd	48	In	49	Sn	50	Sb	51	Te	52	I	53	Xe	54
Lu	71	Hf	72	Ta	73	W	74	Re	75	Os	76	Ir	77	Pt	78	Au	79	Hg	80	Tl	81	Pb	82	Bi	83	Po	84	At	85	Rn	86
Lr	103	Rf	104	Db	105	Sg	106	Bh	107	Hs	108	Mt	109	110	111	112	113	114	115	116	117	118									

B	5	C	6	N	7	O	8	F	9	Ne	10
Al	13	Si	14	P	15	S	16	Cl	17	Ar	18

- 1) Place a large G in boxes that house elements that are gases at SATP, and a large L in boxes that contain elements that are liquids. The rest of the elements are solids; do not write anything in these boxes.
- 2) Darken the "staircase line" that divides metals from non-metals
- 3) Metalloids (i.e. semi-conductors) are elements that have properties midway between metals and non-metals. The metalloids are: B, Si, Ge, As, Sb, Te, Po, At. Identify the metalloids by drawing diagonal lines in these boxes.
- 4) Notice that each group (column) in the periodic table is identified at the top with a number and/or letter. There are two conventions used in the textbook. Label the groups that end in A (e.g. IA, IIA, IIIA, etc).
- 5) Frame and colour the following groups (note: H is not an IA metal): Alkali metals, alkaline earth metals, halogens, noble gasses, transition metals, the inner transition elements (made up of lanthanides and actinides).
- 6) Create a legend for your periodic table.

La	57	Ce	58	Pr	59	Nd	60	Pm	61	Sm	62	Eu	63	Gd	64	Tb	65	Dy	66	Ho	67	Er	68	Tm	69	Yb	70
Ac	89	Th	90	Pa	91	U	92	Np	93	Pu	94	Am	95	Cm	96	Bk	97	Cf	98	Es	99	Fm	100	Md	101	No	102